

Skill Practice 13

Atomic Structure

Name: _____

Date: _____

Hour: _____

- An atom has a mass number of 43 and it also has 21 electrons.
 - How many protons does this atom have?
 - What is the identity of this atom?
 - How many neutrons does this atom have?
- What is an isotope? Give an example.
- A certain *ion* has an atomic number of 16, a mass number of 33, and 18 electrons.
 - What is the charge on the ion?
 - What is the identity of this ion?
 - How many neutrons does the nucleus of this ion have?
- Tritium (an isotope of hydrogen) has 2 neutrons. How many protons does it have? What is its mass number?
- What is the charge on a magnesium ion that has 10 electrons?
- How many neutrons are there in a chromium atom with a mass number of 54?
- Substance E has 29 protons, 28 electrons, and 34 neutrons. Substance F has 29 protons, 27 electrons, and 34 neutrons. Substances E and F can be categorized as...
 A) different elements B) ions C) isotopes D) nuclides E) nucleons
- The element with 38 protons and 45 neutrons could correctly be identified as which element?
- Complete the following table:

Symbol	# of neutrons	# of protons	# of electrons	Atomic #	Mass #
${}_{56}^{136}\text{Ba}^{+2}$					
		25	25		56
	120		79	79	
	21	20	18		